2.Periodic classification of elements -MCQ

Q.1. A) Choose the correct option

1. Doberiener was a …………………….scientist
   - a) German
   - b) American
   - c) Russian
   - d) Indian

2. According to Mendeleev's law the properties of elements are periodic functions of their…………………………..
   - a) Atomic number
   - b) atomic mass
   - c) density
   - d) boiling point

3. Number of electrons in the outermost orbit of the alkali metals is ………….
   - a) 1
   - b) 2
   - c) 3
   - d) 7

4. The valency of the alkali earth metals is 2 which means that their position in the modern periodic table is ………………………………..
   - a) group 2
   - b) group 16
   - c) period 2
   - d) D - group

5. The chemical formula of the chloride of element X is XCl. This compound is a solid having a high melting point. The group of the periodic table of which element x is a member will have ……………. element in the same group
   a. Na
   b. Mg
   c. Al
   d. Si

6. In which block are the non - metals present in the modern periodic table?
   a. S- block
   b. P-block
   c. D-block
   d. F-block

7. Law of octaves was given by …………………..scientist
   a) Doberiener
   b) Newlands
   c) Moseley
   d) Mendeleev

8. . …………...number of elements had been discovered during Newland's time
   a. 56
   b. 65
   c. 63
   d. 36

9. Eka - Boron from Mendeleev's periodic table, was later on called as eka -Boron was later called as ………………..
   a. Gallium
   b. Scandium
   c. Germanium
   d. Molibdenum

Q. 1 B) Choose the odd word out. Also write the reason.

1) Newlands, Moseley, Dobeirner, Mendeleev
2) Fluorine, Sulphur, Bromine, Iodine
3) Sodium, Aluminum, Chlorine, Carbon
4) Nitrogen, Neon, Argon, Helium

Q. 1 C)Complete the pair by identifying the relation

1) Doberiener : triads :: Newlands : ....
2) group 1 : alkali metals :: .... : Halogen
3) solid : Iodine :: .... : Bromine
4) Chlorine : (2,8,7) :: Fluorine : ....

Q. 1. D) Write names

1) Group with zero valency
2) Non- metals in third period
3) two elements having valency
4) Most electronegative atom
5) Atom which is smallest in size
6) Atom with least atomic mass.
7) Most reactive non- metal